



Removal of Cadmium from Aqueous Solutions by a Synthesized Activated Carbon

Mohsen Ansari^{1,2}, Amir Hosein Mahvi³, Mohammad Hossein Salmani², Mohammad Hassan Ehrampoush^{2*}, Mohammad Taghi Ghaneiyani², Mahboubeh Shiraniyan², Parvaneh Tallebi²

¹ Student Research Committee, Shahid Sadoughi University of Medical Sciences, Yazd, Iran.

² Environmental Science and Technology Research Center, Department of Environmental Health Engineering, School of Public Health, Shahid Sadoughi University of Medical Sciences, Yazd, Iran.

³ Solid Waste Research (CSWR) Center, Environmental Research Institute (ERI), Tehran University of Medical Sciences, Tehran, Iran.

ARTICLE INFO

ORIGINAL ARTICLE

Article History:

Received: 5 May 2017

Accepted: 10 July 2018

*Corresponding Author:

Mohammad Hassan Ehrampoush

Email:

ehrampoush@ssu.ac.ir

Tel:

+983538209100

Keywords:

Adsorption,
Activated Carbon,
Solid Waste,
Aqueous Solutions,
Cadmium.

ABSTRACT

Introduction: There are different methods for removal of cadmium from aqueous solutions. Adsorption based methods are among the bests. One of the most important aspects for adsorption techniques is the availability of an accessible and economical adsorbent. This study aims to investigate cadmium removal from aqueous solutions by walnuts shell waste.

Materials and Methods: Walnut Shell (WS) was used as a carbon and was activated by a chemical agent. The surface characteristics of the synthesized Activated Carbon (AC), kinetic and isothermic variation were investigated under laboratory conditions.

Results: The results showed that adsorption could be affected by the amount of adsorbent and pH. The adsorption of cadmium was in the maximum level when pH was 6. Between two investigated adsorption isotherms, Langmuir ($R^2=0.98$) and Freundlich ($R^2=0.97$) isotherms well described the cadmium adsorption on walnut shell produced activated carbon.

Conclusion: Walnut shell as an agricultural waste can be synthesized into a suitable adsorbent to absorb cadmium from aquatic solutions.

Citation: Ansari M, Mahvi AH, Salmani MH, et al. Removal of Cadmium from Aqueous Solutions by a Synthesized Activated Carbon. J Environ Health Sustain Dev. 2018; 3(3): 593-605.

Introduction

In recent decades, increasing demands of industrial products lead to environmental pollution as a critical problem¹. Cadmium is one of the most notorious toxic elements that could be used in many large industries, such as battery manufacturing

industries²; pulp and paper industries³, and etc. The serious adverse health effects of cadmium occur when it is released to the groundwater and other reservoirs⁴. Various natural and manmade release cadmium into the groundwater⁵. The natural processes include soil erosion, drainages from

cadmium contacting soils and atmospheric depositions⁶. The manmade processes include industrial activities such as mining and metallurgical activities, combustion of fossil fuels, application of pesticides, and industrial effluents. Usually, cadmium is observed as Cd⁺² in water resources and serious health effects (e.g. Cancers, Cadmium-induced renal damage, and etc.) occur due to exposure to it⁷. Cadmium as a nephrotoxic element damages kidneys and has adversely effects on bones indirectly or directly⁸. This element may also react with sulfur containing proteins and lead to nervous disorders, pancreas disorders, cardiovascular system disorders, skeletal limb abnormalities, and anemia⁹. Maximum contaminant level (MCL) of cadmium in industrial effluents and drinking water respectively established as 0.25 and 0.005 mg/l by the Iran Environmental Protection Agency (EPA)¹⁰. It is noteworthy that the US EPA also categorized cadmium as a teratogen and a probable human carcinogen¹¹. Common treatment methods for cadmium removal include precipitation¹², membrane filtration¹³, ion exchange,¹⁴ and adsorption¹⁵. Each method has its advantages and disadvantages such as high investment and operating costs, excessive sludge production and inefficiency. On the other hand, the application of recycled adsorbents became a popular method in adsorption process¹⁶. Indeed, adsorption is an effective elimination method; hence activated carbon has been used as an appropriate adsorbent¹⁷. However, due to expensive costs and regeneration difficulties of the adsorbent, the application has been limited significantly in developing countries¹⁸. These limitations induced the researchers and executives for synthesis of applicable, effective, and economical adsorbents from recyclable materials. Over the last few years the application of agricultural wastes, such as date palm seed¹⁹, charcoal ash²⁰, pomegranate peel²¹, beans peel²², poplar tree bark²³, and walnut shell²⁴ for treatment of industrial wastewaters by adsorption of heavy metals was considered as a successful and economical treatment method. These wastes have the suitable characteristics to be used as an adsorbent. Although, they are cheap and ubiquitous,

their disposal need some solid waste management costs²⁵. One of the most popular agricultural wastes for production of activated carbon is walnut shell. For example, in this field of study Jafari Mansuri et.al²⁴, Yu et.al²⁶, Cheng et. al²⁷ and Zafarani et. al²⁸ investigated the use of walnut shell as an adsorbent to remove contaminants from water environment. Walnut shells with high carbon content and low moisture content can be considered as one of the suitable substrates for activated carbon production²⁹. Therefore, considering the limited resources of raw materials and the need for recycling wastes, (in particular waste and agricultural waste, and with an economical, efficient and environmentally friendly approach) this study aimed to investigate the production of activated carbon of walnut shells and its evaluation in removing binary cadmium.

Materials and Methods

Materials

All of the materials used in this study were manufactured by Merck, German corporation.

Absorbent Synthesis

The present study was carried out in a laboratory scale in School of Public Health, Yazd. The walnut shell was prepared from the Shabestar area in East Azarbaijan province, Iran. In order to prepare an adsorbent, first, the walnut shell was crushed by milling and sieved using two sieves of 50 and 40 (microns). Then, the samples were washed twice with distilled water. After drying at ambient temperature, the samples pyrolysis is performed under vacuum in an electric furnace at 350 °C for 2 hours. This product was kept under the name of carbonized walnut shell (CWS). In order to prepare a modified chemical adsorbent, in line with previous studies on the chemical activation of natural adsorbents, the CWS was added to potassium hydroxide solution 0.7 M in the impregnation ratio of KOH to walnut shell (1 g of agent: 1 g of precursor) for 2 hours at 250 °C. The prepared material at this stage was called the Activated Carbon synthesized from Walnut Shell (ACWS). After completion of the chemical activation process, the ACWS was washed several times with distilled

water without ion to remove the excess potassium hydroxide from ACWS.

Adsorbent Characteristics

The ACWS dataset parameters were determined in accordance with the following standards; such as moisture content and apparent density (AWWA standard)³⁰, ash content (ASTM D2866)³¹, and iodine value (ASTM D4607-14)³². The specific surface area and the pore volume of ACWS were also determined by the BET method by the BELSORP Mini II machine. The X-ray powder diffraction (XRD) as a rapid analytical technique was primarily used for the identification phase of the crystalline material and it can provide information on unit cell dimensions. In this study the X-ray diffraction (XRD) was used by the Bruker AXS-D8 Advance device. One of its benefits is the lack of vacuum, which reduces the cost of construction and places it at a higher level than electronic techniques. Furthermore, the XRD is non-contact and non-destructive and requires no hard preparation³³. The pH of the point of zero charge (pH_{zpc}) was measured using the pH drift method. The pH of the sorbent in the diluted water was adjusted between 2 and 12 by adding 0.01 M NaOH and 0.01 M HCl; 0.2 g of the adsorbent was added to 50 cm³ of the solution, and after 24 h, the final pH was measured³⁴.

Kinetic Study

The initial standard stock solutions of Cd(II) ion (about 1000 mg/L) were prepared by dissolution of Cd(NO₃)₂·4H₂O powder (Merck, Germany) in redistilled water. Then various concentrations of cadmium were prepared by diluting the stock solution using double-dissolving water. In the adsorption study of cadmium in the ACWS, a batch equilibration technique was used. Initially, the effect of time on cadmium adsorption in the ACWS was investigated. The initial concentration of cadmium in the single element solution was prepared to be at 100 mg/L, and the ACWS mass were around 0.2 g ± 0.03 g. The optimal pH in case of each metal ion was adjusted to 6. The

kinetics solution was stirred on a shaker at room temperature 22 ± 2 °C at 120 rpm constant speed. The mixtures were filtered with 0.45 μm filters, and then the filtrates residue was measured using a spectrophotometer at a wavelength of 545 nm [34]. The kinetics data were optimized using Lagergren pseudo-first-order equation and pseudo-second-order equation. Finally, in this study the influencing factors on the cadmium adsorption such as pH, initial concentration of cadmium, and adsorbent dose were investigated.

Isotherms Study

The adsorption ability of cadmium on the ACWS was obtained according to the initial runs of adsorption isotherm. The adsorption isotherm was carried out at the initial pH of 6. During experiments, the cadmium concentration from 5 to 350 mg/L range was used. The adsorption analysis was consisting of 50 mL of solution and 0.2 g ± 0.03 g of the ACWS. Adsorption isotherms solution was stirred on a shaker with the temperature of 22 ± 2 °C and the equilibrium time of 24 h. In this study, Freundlich and Langmuir were used to fit the sorption isotherms.

The Langmuir absorption model is shown in the following equation:

$$q_e = \frac{q_m K_L C_e}{1 + K_L C_e}$$

In this equation, q_e is the amount of dye adsorbed per unit mass of adsorbent at equilibrium conditions (mg/g), C_e is the dye equilibrium concentration (mg/L), q_m is the monolayer capacity of the adsorbent (mg/g), and K_L is the Langmuir constant.

The Freundlich isotherm describes the adsorption on heterogeneous surfaces with different energy of adsorption onto the adsorbent surfaces sites and is expressed by the following equation:

$$q_e = K_f C_e^{1/n}$$

In this equation, K_f (mg/g).(L/mg)^{1/n} and n are the Freundlich coefficients, K_f is related to adsorption capacity and n indicates the effect of concentration on the adsorption capacity.

Results

Absorbent Characteristics

Characteristics of activated carbon prepared from walnut shell are described in Table 1. As it can be seen, apparent density, pH zero point of charge, ash content, iodine number, specific surface area, and pore volume of the ACWS were 0.78 (g/L), 7.3, 2.3 (%), 625 (mg/g), 1477 (m²/g), 1.3 (cm³/g), respectively.

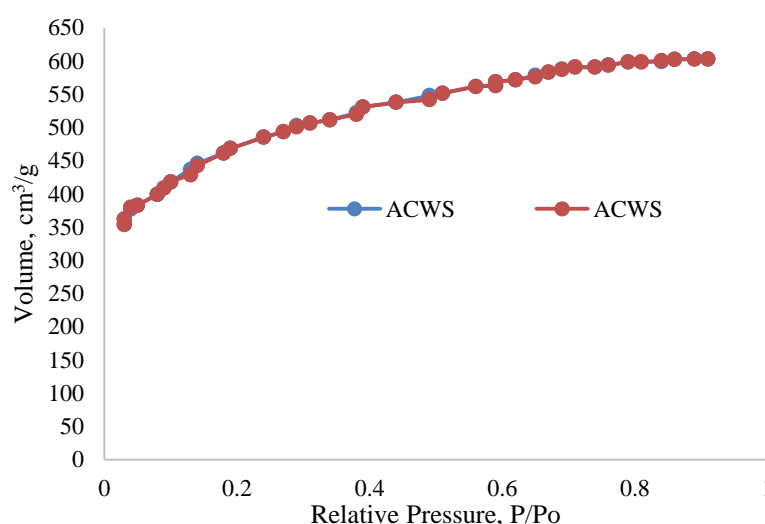
The adsorption/desorption isotherms of the CWS and ACWS by the KOH are presented in Figure 1. The use of activating agent has a significant effect on the shape of the N₂ adsorption/desorption isotherms. Furthermore, Figure 2 illustrates the pore size distribution of the

cavity structure of the ACWS and CWS. According to Figure 2, the pore size distribution of the ACWS and CWS are in the range of 2 nm to 4 nm in the mesoporous scale.

The XRD spectrum of carbonized walnut shell before and after activation is presented in Figure 3. This figure indicates that there was a tetragonal phase only for the ACWS. However, after the activation of carbonized walnut shell, the ACWS presented only the crystallinity phase, as shown in the diffraction peaks at 43.5°, 52.3°, and 72.8°. Moreover, there are diffraction peaks at 25° and 43° (101), which are assigned to the carbon phase as the graphite structure in the ACWS.

Table 1: Characteristics of activated carbon prepared from walnut shell

Characteristics	Unit	Value
Apparent density	g/L	0.78
pH _{zero point of charge}	-	7.3
Ash content	%	2.3
Iodine number	mg/g	625
Specific surface area BET	m ² /g	1604
Microporous	cm ³ /g	0.684
Mesoporous	cm ³ /g	0.196
Total volume	cm ³ /g	0.880
Diameter	(nm)	2.19



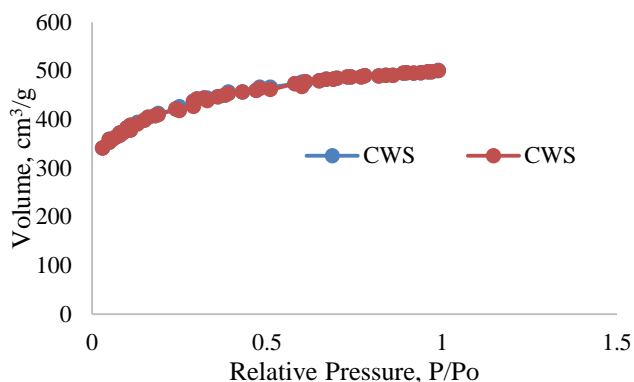


Figure 1: Adsorption and desorption isotherms N₂ at 77 °K of ACWS and CWS

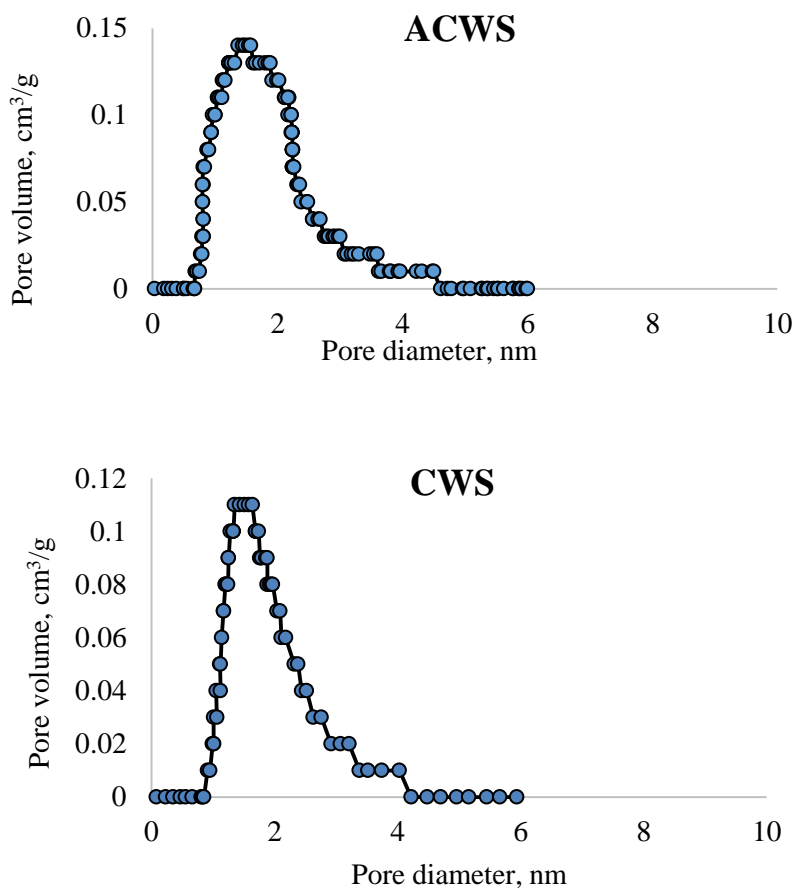


Figure 2: Pore size distribution of CWS and ACWS

Sorption Kinetic

The adsorption of cadmium on the ACWS was studied as a function of time, and the results are

shown in Figure 3. The sorption kinetic analysis revealed that the percentage of Cd (II) adsorption increased with time (Figure 4).

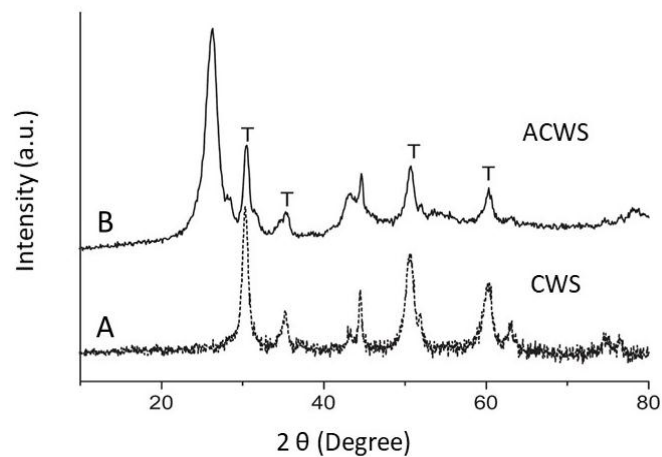


Figure 3: XRD patterns of activated carbon synthesized from walnut shell

Figure 4 indicates that the adsorption process of cadmium on the ACWS had two phases, which is in line with the previous research studies³⁵. The first rapid step is attributed to the surface

adsorption of cadmium on the ACWS. The following slow step is explained that the adsorbed cadmium was distributed from the outer to inner surfaces of ACWS³⁵.

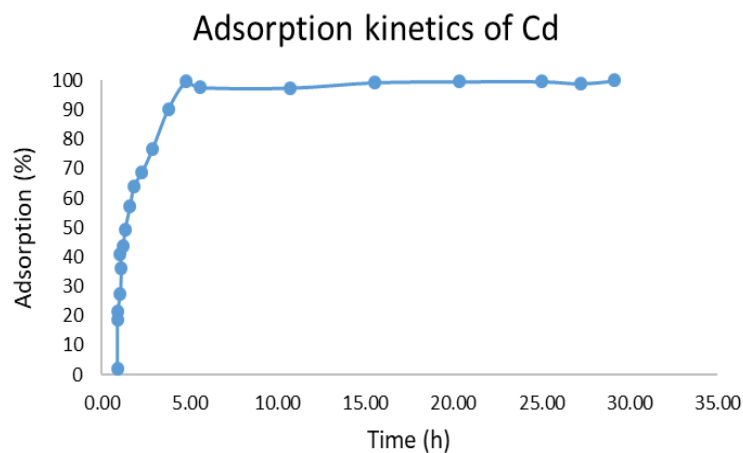


Figure 4: Adsorption kinetics of Cd (II) ions onto ACWS, $m = 0.2$ g, $V = 50$ mL, $C = 100$ mg/L, $\text{pH} = 6$, $T = 22 \pm 2$ °C.

Effect of pH

The pH of the solution has some function such as determining the surface charge of the adsorbent, the degree of ionization and adsorption formation¹⁷. The effect of pH on cadmium adsorption by the

ACWS is shown in Figure 5. This figure represents that the minimum and the maximum absorption levels of cadmium are observed at pH 2 and pH 6, respectively.

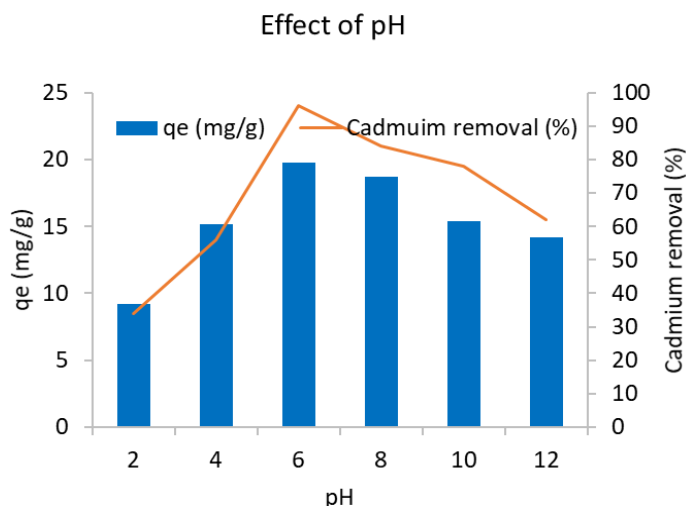


Figure 5: Effect of pH on the adsorption capacity and removal percentage of cadmium onto ACWS (dose of activated carbon = 0.5 g/L, Cd(II) = 25 mg/L, temperature = 22 °C)

Effect of Adsorbent Dose

The adsorption of cadmium as a function of adsorbent dose is shown in Figure 6. The experiment was carried out at V = 50 mL, Cd (II) = 25 mg/L, pH = 6, T = 22 ± 2 °C. The cadmium removal percentage increased significantly from 36.56 to 98.21 % when the adsorbent dose was increased from 0.25 to 1.25

g/L that was due to the substantial increase in available adsorption sites on the ACWS. As Figure 1 indicates, the adsorption capacity of cadmium was decreased from 19 to 9.2 mg/g with increasing the amount of the ACWS. A similar behavior was reported for the adsorption of Cr (VI) and Fe (II) from aqueous solution on walnut shell³⁶.

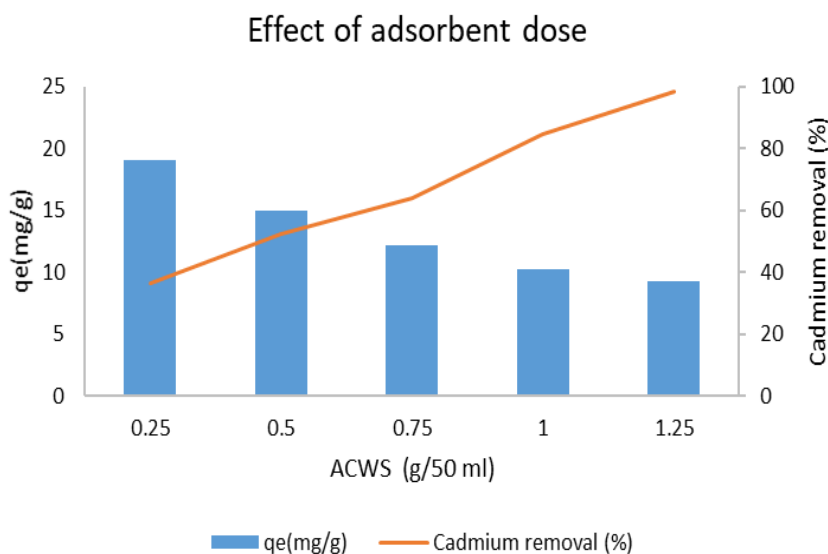


Figure 6: Effect of adsorbent dose on the adsorption capacity and removal percentage of cadmium onto ACWS (pH 6, temperature 22 °C Cd (II) = 25 mg/L)

Adsorption Isotherms

Langmuir Isotherm

The adsorption isotherm is one of the important factors used in the design of absorption systems. In

fact, the adsorption isotherm is explained as an interaction between the absorbent and the absorbed object. Therefore, it is considered as an essential factor in determining the absorbent capacity and

optimizing the adsorption capacity. In the Langmuir isotherm, it is assumed that the single-layer adsorption is occurred in homogeneous sites on the adsorbent, in contrast to the Freundlich isotherm to describe heterogeneous systems.

The adsorption equilibrium data fitted with Langmuir and Freundlich models are shown in Figure 7. This figure shows a steep slope that specified the high efficiency of adsorbent at a low cadmium concentration and the slope decreased with increasing cadmium concentration. Furthermore, Constant parameters and correlation

coefficients calculated for Langmuir and Freundlich models at different initial cadmium concentration for the ACWS are shown in Table 2. The determination coefficient (R^2) for cadmium adsorption was equal to 0.98, which indicated a good fit of Langmuir single-layer model for cadmium adsorption by the ACWS. In other words, the representation of the mathematical adaptation is very good, which may be due to the explanation of similar adsorption sites on the surface of the ACWS. Since Langmuir isotherm assumes the absorbing surface homogeneously.

Table 2: Isotherms parameters and determination coefficients (R^2)

Freundlich model			Langmuir model		
R^2	K_F	$1/n$	R^2	$K_L(L/mg)$	$q_m(mg/g)$
0.97	23.63	0.19	0.98	0.072	68.31

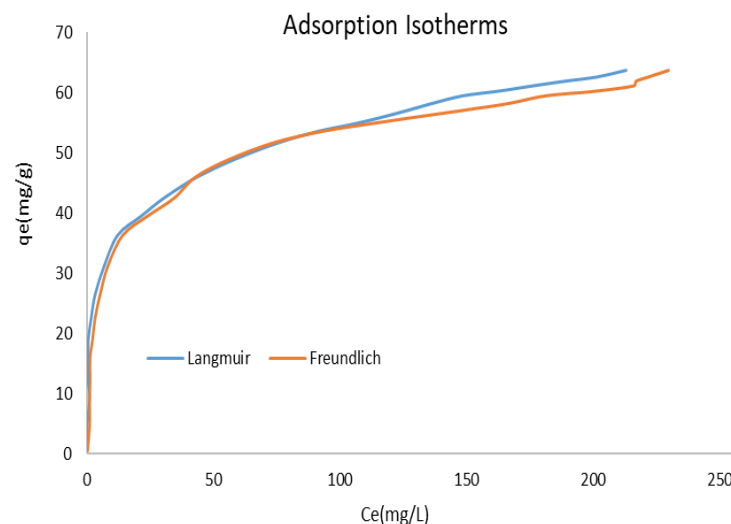


Figure 7: Freundlich and Langmuir isotherms for cadmium adsorption onto ACWS at different initial cadmium concentration (3 to 230 mg/l) (pH 6, temperature = 22 °C, dose of ACWS = 2 g/L)

Freundlich Isotherm

The results of the Freundlich model (Figure 7) showed that similar to the Langmuir slope, the slope of Freundlich model was sharp and showed the existence of high efficiency of the ACWS at a low cadmium concentration and the slope decreased with increasing cadmium concentration. As it is revealed in Table 2, R^2 for cadmium adsorption in the Freundlich model was equal to 0.97, which indicated a good fit of Freundlich model for cadmium adsorption by the ACWS.

Discussion

As a result, above the pH_{zpc} of 7.3, the ACWS has a negative charge and is capable of attracting cations. The quality of the ACWS is defined by various properties such as moisture content, ash content, and iodine number. In this study, the values related to moisture content, ash content, and iodine number were in the ranges that a high quality activated carbon needs them. The both isotherms of the CWS and ACWS are of type I in the IUPAC system. Based on the IUPAC

classification, the porous media structure can contain holes smaller than 2 nm, called micro hollow (based on the average dimensions of the holes). Holes between 2 to 50 nm, are called meso holes, and holes larger than 50 nm, are called Macro³⁷. This type of isotherm, often called Langmuir, is rarely seen in non-porous materials, and is suitable for compounds with high-density cavities (less than 2 nm). This indicates that the pore structure of samples is mostly microporous. As shown in Figure 1, the BET surface area of the ACWS was higher than that of CWS. This is related to the development of the high micropore volume and control of pore-size distributions due to use of activating agent. The type I of the IUPAC system indicates that the samples exhibit a pore size distribution on the micropore range which was confirmed by the micropore volumes in Table 1. As Figure 2 reveals, the pore volume (cm³/g) in the ACWS was more than CWS. The results indicate that the ACWS is ideal for creating active sites to absorb multi-capacity metal contaminants. Therefore, the effect of the KOH was associated with stronger release of volatiles in the activation process and conversion to microspores³⁸. Moreover, adding the KOH may prevent the formation of tar known as the pore-blocking agent; therefore, it can increase the surface area. This finding is in line with the studies of Zhanyong et al.²³, Lewicka³⁹, and Qiongfeng et al.⁴⁰ who used the KOH as an

activating agent.

The physical and chemical characteristics of a substance are related to the sorption kinetics, which also influence the adsorption process. As shown in Figure 3, the elemental composition of the ACWS showed a carbon phase as the graphite structure in the ACWS. Therefore, the surface of the ACWS can show a greater amount of polar groups responsible for binding the heavy metal ions⁴¹.

Accordingly, Figure 4 shows that the adsorption kinetics of Cd (II) was slow and to reach the equilibrium state, the required contact time with the ACWS surfaces should be 6 h. According to pseudo-second-order data ($R^2 = 0.98$), the kinetic fitting plots were used to explain the time base mechanisms of adsorption process (Figure 8). Table 3 represents the kinetic parameters obtained for cadmium sorption kinetics. It can be seen that only the pseudo-second-order model provided the best fitting for all of the experimental data.

The regression coefficient can be extracted from Figure 4. The calculated regression coefficient show higher than 0.98 cadmium adsorption onto ACWS (Table 3). It was suggested (Kołodyńska et al., 2012) that the kinetics of metal ions following the pseudo-second-order model is controlled by a chemisorption process. The obtained second-order sorption kinetic is in agreement with other authors' findings⁴²⁻⁴⁴.

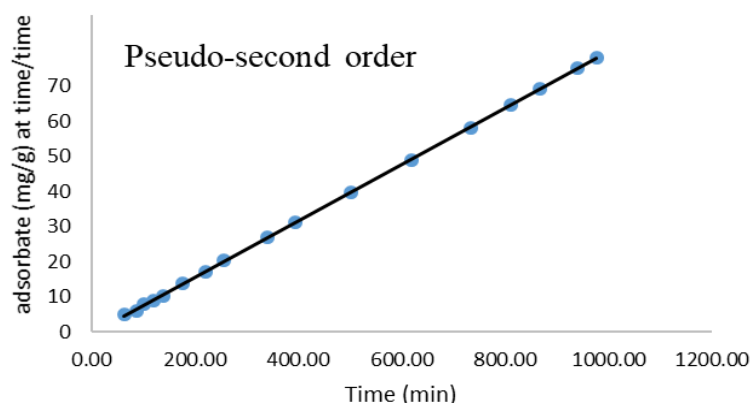


Figure 8: The kinetic fitting plots according to pseudo second order for ACWS

Table 3: Kinetic models

Kinetic model					
Pseudo-first order		Pseudo-second order			ak
k_1	R^2	k_2	R^2	a_{eq}	
0.0036	0.671	0.06	0.98	20.61	22.89

The result of pH effects (Figure 5) was consistent with Elouear et al. that found the maximum adsorption of nickel and cadmium was at pH 6 and the metal ions removal was nearly constant for $pH > 6$ ⁴⁵.

At lower pH_s , H_3O^+ ions compete with metal ions to the adsorption sites; therefore, absorbing metal ions increases with increasing pH from 2 to 6⁴⁶. The value of q decreases at high pH_s due to metal ions deterioration. Previous studies have shown that optimized pH values have been obtained for adsorption of metal ions from aqueous solutions under acidic conditions^{47, 48}. In another study by Veemaraj Kannan, the absorption of cadmium bivalent capacity ions by activated carbon synthesized from the rubber tree in a pH range of 2 to 10 was studied. The results of their studies showed that by increasing the pH of the aqueous solution, the percentage of cadmium removal would increase and pH of about 7 would be the most suitable pH for the adsorption of cadmium by this natural adsorbent⁴⁹. The minimum metal adsorption at low pHs may be due to the high concentration and mobility of H^+ ions, which are preferably absorbed instead of metal ions⁵⁰. At high pHs, cadmium adsorption increases as a result of the reduction in the number of H^+ ions and the increase of negative charge ligands⁵¹. At pH below the pH_{zpc} , the surface charge is absorbent; therefore, this positive charge on the ACWS is suitable for adsorption of bivalent cadmium cations. At pH higher than the pH_{zpc} , pure charge on the surface of the ACWS in this study was negated, resulting in the absorption of bivalent cadmium cations. Meanwhile, at higher pH, hydroxyl ions increase in the solution and compete with cadmium ions to adsorb onto active sites of the ACWS and cadmium absorption decreases. Since the amount of

adsorbent pH_{zpc} used is 7.7, the reduction in adsorption efficiency in pH conditions above 7 is reasonable. Studies done by other researchers have shown that optimum pH for bivalent cadmium is between 5 and 7.

According to Table 2, the monolayer capacity of the adsorbent (q_m) for the ACWS was obtained 68.31 mg / g. Various studies presented the maximum absorption capacity of cadmium by other adsorbents. For example, Zubrik et al. calculated the maximum absorption capacity 107 mg of cadmium per gram of the activated carbon from agricultural waste⁵². In another study, Taghavi et al. investigated the efficiency of crystalline nanotubes carbon to remove component cadmium and showed a well-fitting Langmuir model ($R^2 = 0.994$) with q_m of 43.47 mg/g⁵³. However, it should be noted that direct comparison of various studies is difficult because the experimental conditions are very different.

Conclusion

This study was conducted to produce a valuable adsorbent of agricultural waste. Walnut shell has properties such as high carbon content, low humidity, availability, and successfully used as a cadmium adsorbent. Preparing activated carbon with less heat energy usage can be considered as a substitute for industrial activated carbon. The activated carbon derived from walnut shell was significantly effective in removing cadmium from the aquatic environment.

Funding

This study was supported by the Student Research Committee of Shahid Sadoughi University of Medical Sciences with grant number 5088.

Conflict of Interest

The authors declare that there is no conflict of interest.

This is an Open Access article distributed in accordance with the terms of the Creative Commons Attribution (CC BY 4.0) license, which permits others to distribute, remix, adapt and build upon this work, for commercial use.

References

- Ahmadi E, Yousefzadeh S, Ansari M, et al. Performance, kinetic, and biodegradation pathway evaluation of anaerobic fixed film fixed bed reactor in removing phthalic acid esters from wastewater. *Health Sci Rep.* 2017; 7: 41020.
- Addala A, Belattar N. Adsorption of Cd and Pb metal ions onto chelating resin and their application in removal of lead from battery factory wastewaters. *Indian Journal of Chemical Technology.* 2017; 24(6): 601-7.
- Pokhrel D, Viraraghavan T. Treatment of pulp and paper mill wastewater—a review. *Sci Total Environ.* 2004; 333(1-3): 37-58.
- Wang L, Cui X, Cheng H, et al. A review of soil cadmium contamination in China including a health risk assessment. *Environ Sci Pollut Res Int.* 2015; 22(21): 16441-52.
- Pinot F, Kreps SE, Bachelet M, et al. Cadmium in the environment: sources, mechanisms of biotoxicity, and biomarkers. *Rev Environ Health.* 2000; 15(3): 299-324.
- Yadanaparathi SKR, Graybill D, von Wandruszka R. Adsorbents for the removal of arsenic, cadmium, and lead from contaminated waters. *J Hazard Mater.* 2009; 171(1): 1-15.
- Thornton I. Sources and pathways of cadmium in the environment. *IARC Sci Publ.* 1992; (118): 149-62.
- Jarup L, Akesson A. Current status of cadmium as an environmental health problem. *Toxicol Appl Pharmacol.* 2009; 238(3): 201-8.
- Nordberg GF. Historical perspectives on cadmium toxicology. *Toxicol Appl Pharmacol.* 2009; 238(3): 192-200.
- Agency TIEP. Sewage outlet standard. Tehran: IEPA; 1994.
- EPA. Aquatic Life Ambient Water Quality Criteria Cadmium-2016. Office of Water; 2016. Contract No.: 820-R-16-002.
- Tan H, He W, Han H, et al. Cadmium removal in water emergency treatment of conventional process enhanced by chemical precipitation. *Huan Jing Ke Xue.* 2013; 7(3): 848-852.
- Ennigrou DJ, Sik Ali MB, Dhahbi M, et al. Cadmium and zinc removal from water by polyelectrolyte enhanced ultrafiltration. *Membr Water Treat.* 2014; 5(3): 183-195.
- Haklits I, Nagy A, Kiss L, et al. Cadmium removal from waste water by ion exchange. *Reactive Polymers, Ion Exchangers, Sorbents.* 1986; 7(2-3): 321.
- Wang Z, Huang G, An C, et al. Removal of copper, zinc and cadmium ions through adsorption on water-quenched blast furnace slag. *Desalination and Water Treatment.* 2016; 57(47): 22493-506.
- Bogusz A, Oleszczuk P, Dobrowolski R. Application of laboratory prepared and commercially available biochars to adsorption of cadmium, copper and zinc ions from water. *Bioresour Technol.* 2015; 196.
- Fu F, Wang Q. Removal of heavy metal ions from wastewaters: a review. *J Environ Manage.* 2011; 92(3): 407-18.
- Foo K, Hameed B. Utilization of rice husk ash as novel adsorbent: a judicious recycling of the colloidal agricultural waste. *Adv Colloid Interface Sci.* 2009; 152(1-2): 39-47.
- Afdhol M, Amiliana R, Hanafi A, editors. Preparation of activated carbon from palm shells using KOH and ZnCl₂ as the activating agent. *IOP Conference Series: Earth and Environmental Science*; 2017: IOP Publishing.
- Javadian H, Ghorbani F, Tayebi H, et al. Study of the adsorption of Cd (II) from aqueous solution using zeolite-based geopolymer, synthesized from coal fly ash; kinetic, isotherm and thermodynamic studies. *Arabian Journal of Chemistry.* 2015; 8(6): 837-49.

21. Senthilkumar T, Chattopadhyay S, Miranda LR. optimization of activated carbon preparation from pomegranate peel (*Punica granatum* Peel) using RSM. *Chem Eng Commun.* 2017; 204(2): 238-48.
22. Plaza-Recobert M, Trautwein G, Perez-Cadenas M, et al. Preparation of binderless activated carbon monoliths from cocoa bean husk. *Microporous Mesoporous Mater.* 2017; 243: 28-38.
23. Li Z, Gao X, Wu L, et al. Preparation of activated carbons from poplar wood by chemical activation with KOH. *J Porous Mater.* 2017; 24(1): 193-202.
24. Jafari-Mansoorian H, Farzadkia M, Ansari M, et al. Evaluating the activated carbon prepared from walnut in removal of arsenic from aqueous solution. *Journal of Safety Promotion and Injury Prevention.* 2016; 3(4): 287-94.
25. Sud D, Mahajan G, Kaur M. Agricultural waste material as potential adsorbent for sequestering heavy metal ions from aqueous solutions—A review. *Bioresour Technol.* 2008; 99(14): 6017-27.
26. Yu Q, Li M, Ji X, et al. Characterization and methanol adsorption of walnut-shell activated carbon prepared by KOH activation. *Journal of Wuhan University of Technology-Mater Sci Ed.* 2016; 31(2): 260-8.
27. Cheng L, Sun L, Xue W, et al. Adsorption equilibrium and kinetics of Pb (II) from aqueous solution by modified walnut shell. *Environ Prog Sustain Energy.* 2016; 35(6): 1724-31.
28. Zafarani HR, Bahrololoom ME, Noubactep C, et al. Green walnut shell as a new material for removal of Cr (VI) ions from aqueous solutions. *Desalination and Water Treatment.* 2015; 55(2): 431-9.
29. Zabihi M, Asl AH, Ahmadpour A. Studies on adsorption of mercury from aqueous solution on activated carbons prepared from walnut shell. *J Hazard Mater.* 2010; 174(1-3): 251-6.
30. AWWA. B604-96: AWWA Standard for Granular Activated Carbon. Denver, CO 80235 USA: AWWA; 1997.
31. American Society for Testing and Materials. Standard Test Method for Total Ash Content of Activated Carbon. West Conshohocken, PA : ASTM, 1999.
32. ASTM D4607-14, Standard Test Method for Determination of Iodine Number of Activated Carbon, ASTM International, West Conshohocken, PA, 2014.
33. Cullity BD. *Elements of X-ray Diffraction*: Addison-Wesley Publishing Company; 1978.
34. Kołodynska D, Gęca M, Pylypchuk IV, et al. Development of new effective sorbents based on nanomagnetite. *Nanoscale Res Lett.* 2016; 11: 152.
35. Thirumavalavan M, Lai YL, Lee JF. Fourier transform infrared spectroscopic analysis of fruit peels before and after the adsorption of heavy metal ions from aqueous solution. *J Chem Eng Data.* 2011; 56(5): 2249-55.
36. Ghasemi M, Ghoreyshi A, Younesi H, et al. Synthesis of a high characteristics activated carbon from walnut shell for the removal of Cr (VI) and Fe (II) from aqueous solution: single and binary solutes adsorption. *Iranian Journal of Chemical Engineering.* 2015; 12(4): 28-51.
37. Sing KSW. Reporting physisorption data for gas/solid systems with special reference to the determination of surface area and porosity (Recommendations 1984). *Pure and applied chemistry.* 1985; 57(4): 603-19.
38. Yang J, Qiu K. Preparation of activated carbons from walnut shells via vacuum chemical activation and their application for methylene blue removal. *Chem Eng J.* 2010; 165(1): 209-17.
39. Lewicka K. Activated carbons prepared from hazelnut shells, walnut shells and peanut shells for high CO₂ adsorption. *Pol J Chem Technol.* 2017; 19(2): 38-43.
40. Yu Q, Li M, Ning P, et al. Preparation and phosphine adsorption of activated carbon prepared from walnut shells by KOH chemical activation. *Sep Sci Technol.* 2014; 49(15): 2366-75.
41. Ngah WSW, Hanafiah M. Removal of heavy metal ions from wastewater by chemically

- modified plant wastes as adsorbents: a review. *Bioresour Technol.* 2008; 99(10): 3935-48.
42. Dubey A, Mishra A, Singhal S. Application of dried plant biomass as novel low-cost adsorbent for removal of cadmium from aqueous solution. *Int J Environ Sci Technol.* 2014; 11(4): 1043-50.
43. Huggins TM, Haeger A, Biffinger JC, et al. Granular biochar compared with activated carbon for wastewater treatment and resource recovery. *Water Res.* 2016; 94: 225-32.
44. Tang N, Niu CG, Li XT, et al. Efficient removal of Cd²⁺ and Pb²⁺ from aqueous solution with amino- and thiol-functionalized activated carbon: Isotherm and kinetics modeling. *Sci Total Environ.* 2018; 635: 1331-44.
45. Elouear Z, Bouzid J, Boujelben N. Removal of nickel and cadmium from aqueous solutions by sewage sludge ash: study in single and binary systems. *Environmental technology.* 2009; 30(6): 561-70.
46. Reddad Z, Gerente C, Andres Y, et al. Adsorption of several metal ions onto a low-cost biosorbent: Kinetic and equilibrium studies. *Environ Sci Technol.* 2002; 36(9): 2067-73.
47. Kazemipour M, Ansari M, Tajrobehkar S, et al. Removal of lead, cadmium, zinc, and copper from industrial wastewater by carbon developed from walnut, hazelnut, almond, pistachio shell, and apricot stone. *J Hazard Mater.* 2008; 150(2): 322-7.
48. Wu P, Wu W, Li S, et al. Removal of Cd²⁺ from aqueous solution by adsorption using Fe-montmorillonite. *J Hazard Mater.* 2009; 169(1-3):824-30.
49. Kannan N, Veemaraj T. Detoxification of toxic metal ions by sorption onto activated carbon from Hevea brasiliensis bark. *Global NEST journal.* 2010; 2(12): 197-205.
50. Ajmal M, Rao RAK, Ahmad R, et al. Adsorption studies on Citrus reticulata (fruit peel of orange): removal and recovery of Ni (II) from electroplating wastewater. *J Hazard Mater.* 2000; 79(1-2): 117-31.
51. Norton L, Baskaran K, McKenzie T. Biosorption of zinc from aqueous solutions using biosolids. *J Environ Manage.* 2004; 8(3-4): 629-35.
52. Zubrik A, Matik M, Hredzák S, et al. Preparation of chemically activated carbon from waste biomass by single-stage and two-stage pyrolysis. *J Clean Prod.* 2017; 143: 643-53.
53. Taghavi M, Zazouli MA, Yousefi Z, et al. Kinetic and isotherm modeling of Cd (II) adsorption by L-cysteine functionalized multi-walled carbon nanotubes as adsorbent. *Environ Monit Assess.* 2015; 187(11): 682.